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C: When  $n = 3$ ,  $l = 0, 1, \text{ or } 2$ . The highest value for  $l$  in this case is 2, which corresponds to the  $d$  subshell. Although the  $3d$  block appears to be part of the fourth period, it still has the principal quantum number  $n = 3$ . In general, the subshells within an energy shell increase in energy as follows:  $s < p < d < f$  (although there is no  $3f$  subshell).

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