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in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

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the following questions in the space provided. 1. Given the following

equation: $C_3H_4(g) + x \cdot O_2(g) \rightarrow 3CO_2(g)$

+ $2H_2O(g)$ a. What is the value of the

coefficient . x. in this equation? b. What

is the molar mass of C_3H_4 ? c. How many

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moles are in an 8.0 g sample of C_3H_4 ? 2.
a. What is meant by . ideal conditions

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